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Articular cartilage: Can Articular Cartilage be considered a Smart Material?

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Abstract

It has been shown that the "smart surface" of cartilage is highly hydrophilic when wet and hydrophobic when air-dry. The minimum surface energy of the bilayer was found to spread over a wide range of pH (6.5 to 9.5). The decreased number of bilayers changed the wettability and lowered the PL lubricant properties. In conclusion, a cartilage smart material can significantly react to changes pH and moisture.

Key Words: articular cartilage; hydrophilic; hydrophobic; pH; wettability

The biological tissue of the articular cartilage (AC) is amphoteric in its natural condition transforming surface from the hydrophilic to hydrophobic in air-dry condition (Fig.1 a, b). The profiles of the interfacial surface energy ($^{\gamma}$) of spherical lipid bilayer model "bell-shaped curve" vs. pH change the surface charge from positive, neutral to negative, the articular cartilage can be named a "smart material" (**Fig. 1c**). The understanding of AC amphoteric surface will be essential for understanding the joints lubrication mechanism, the effective repair and regeneration of the degraded knee joint.

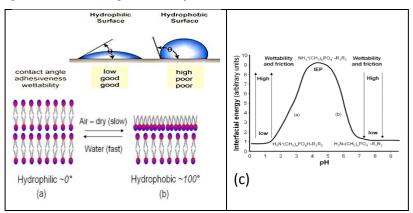


Figure 1: (a, b) The smart-surface constitution of superficial phospholipid bilayer of articular cartilage in (a) wet (hydrophilic) and (b) at air-dry condition (hydrophobic). Fig. 1 (c) The profiles of the interfacial surface energy ($^{\gamma}$) of spherical lipid bilayer model "bell-shaped curve" vs. pH of buffer solution.

(a) (pH 0.2 to 4.0 (-NH₃⁺ \rightarrow -NH_{2.}). (b) (pH 4.0 to 6.5) (-PO₄H \rightarrow -PO₄⁻) IEP : H₂N (CH₂)_n PO₄H-R₁R₂ \leftrightarrows H₃N⁺(CH₂)_n PO₄⁻-R₁R₂

The isoelectric point, IEP, is at a pH of ~ 4 when a PL molecule carries no net electrical charge H_2N (CH_2)_n $PO_4H-R_1R_2 \leftrightarrows H_3N^+$ (CH_2)_n $PO_4^--R_1R_2$. Changes of the interfacial energy correspond to the amino ($-NH_3^+ \rightarrow -NH_2$) transition with a low pH, after IEP the phosphate ($-PO_4H \rightarrow -PO_4^-$) transition at a higher pH range of 4 to 6.5. The isoelectric point, IEP, with the pH at which a phospholipid molecule carries no net electrical charge.

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The wetted surfaces (pH~6.5 to 9.5) of the phospholipid membranes are negatively charged (-PO₄-). wettability can be measured relative to the differences in the charge density of the functional phosphate (-PO₄-) group. In this regard, the wettability of a hydrated surface is characterized by the concentration of charged anionic phosphate (-PO₄-) groups that are deactivated when the surface is dehydrated. The dehydration of the phospholipid bilayer surface activates hydrophobic groups, R (CH₂)_n - due to the formation of a hydrophobic monolayer, Fig.1. The hydrophilic surface in the course of dehydration, for example, via air-drying, undergoes a slow increase in the wettability contact angle indicating conformational changes in the surface (flip-flop) phospholipid molecules. Poor lubrication in animal joints, particularly on the articular surface of cartilage, can be attributed to deterioration of the bilayer surface, where the wettability or contact angle (θ) changes from 104° to less than 70° (Wettability was measured in air-dry condition)

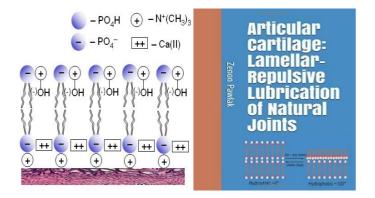


Figure 2: (A) The hydrophilic model of cartilage surface (B) shows the cartilage transformation of the hydrophilic bilayer at wet condition to hydrophobic monolayer (at air-dry room condition).

The lamellar-repulsive mechanism is supported by phospholipid lamellar phases and charged macromolecules from synovial fluid between negatively charged cartilage surfaces. The low friction results as a consequence of a short-range hydration repulsion (nanometer-scale) between the interface of negatively charged (-PO₄-) cartilage surfaces and contribution of macromolecules in the synovial fluid and most importantly a lamellar slippage of bilayers. The acid-base properties of the phospholipid bilayer significantly influence their wettability and surface friction properties Fig.2.

In conclusion, a cartilage smart material can significantly react to changes pH and moisture.

References

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